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Two equilibria must

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$\text{Cu (s)} + \text{H}_2\text{(g)} \rightleftharpoons \text{H}_2\text{O (l)}$
+ Cu at equilibrium
 $\Delta G^\circ = -nFE^\circ = -2 \times 96485 \times 0.34 = -65810 \text{ J}$
+ $\frac{1}{2} \text{O}_2\text{(g)} \rightleftharpoons \text{CuO}$ 1 at
equilibrium $\Delta G^\circ = -nFE^\circ = -2 \times 96485 \times 0.15 = -28945 \text{ J}$
In the final state,
only O_2 is present.
Therefore $\Delta G = -nFE^\circ = -2 \times 96485 \times 0.15 = -28945 \text{ J}$.
In an excess of O_2 , all
the 1 copper is
oxidized. Therefore
 $\Delta G = 0$ or $\Delta G = -nFE^\circ = -2 \times 96485 \times 0.15 = -28945 \text{ J}$.

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