

# Access Free Mixed Gas Laws Answer Key

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The Gas Laws

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Bookmark File PDF Mixed Gas Laws Answer Key WS 5.5: Mixed Gas Law Problems. Directions: Solve the following problems. Round your answers using significant figures. 1) Calculate the mass of 15.0 L of  $\text{NH}_3$  at  $27^\circ \text{C}$  and 900.0 mm Hg. 2) A volume of 26.5 mL of nitrogen gas was collected in a tube at a temperature of  $17^\circ \text{C}$  and a pressure of 737 ...

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Mixed Gas Laws Worksheet Answers Mixed Gas Laws Worksheet - Solutions 1) How many moles of gas occupy 98 L at a pressure of 2.8 atmospheres and a temperature of 292 K?  $n = \frac{PV}{RT} = \frac{(2.8 \text{ atm})(98 \text{ L})}{(0.0821 \text{ L}\cdot\text{atm}/\text{mol}\cdot\text{K})(292 \text{ K})} = 11$  moles of gas 2) If 5.0 moles of  $\text{O}_2$  and 3.0 moles of  $\text{N}_2$  are placed in a 30.0 L tank at a temperature of 25 °C

Mixed Gas Laws Worksheet Answers

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Answer Key Keywords: gas, laws, mixed, practice, answer, key Created Date: 11/5/2020 7:19:20 AM

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The Mixed Gas Laws Worksheet Answers will explain the following: how much is in a tank, what type of fuel is used, how and where it is stored, and when it is available to be used. The answers that you will receive for these questions will vary from one state to another, but you will most likely receive similar answers.

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atm)(98 L) = 11 moles of gas RT (0.0821 L.atm/mol.K)(292 K) 2) If 5.0 moles of O<sub>2</sub> and 3.0 moles of N<sub>2</sub> are

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