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SHORT ANSWER Answer the following questions in the space provided. 1. Explain the difference between the mass numberand the atomic numberof a nuclide. Mass number is the total number of protons and neutrons in the nucleus of an isotope. Atomic number is the total number of protons in the nucleus of each atom of an element. 2.

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**1 Matter and Change - HUBBARD'S CHEMISTRY**

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**Study GuideChapter 5-21 Answer Key**

SHORT ANSWER Answer the following questions in the space provided. 1. Determine whether each of the following is an example of observation and data,a theory, a hypothesis, a control, or a model. observation and data a. A research team records the rainfall in inches per day in a prescribed area of the rain forest. The

**2 Measurements and Calculations**

phase energy (Eph) -energy stored in system due to arrangement of particles that exert attractions on one another. (energy account involved when phase changes occur) -Attractions result in decrease of energy of a system of particles. -As particles become more tightly bound, their phase energy is lowered. -Solids possess lowest phase energy, and liquids possess more because the particles move freer.

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I ordered two books for my high schooler's Chemistry class, "section reviews" and "study guide" by the same authors, to go with her text book "Modern Chemistry". I was very disappointed to find out that these two books are identical in the first 80 pages, the same title, the same contents, the same page number etc, but the cover is completely ...

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SHORT ANSWER Answer the following questions in the space provided. 1. In cathode-ray tubes, the cathode ray is emitted from the negative electrode, which is called the cathode . 2. The smallest unit of an element that can exist either alone or in molecules containing the same or different elements is the .atom 3.

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the intermolecular attractions resulting from the constant motion of electrons and the creation of instantaneous dipoles; forces act between all atoms AND molecules; forces are NOT present in intramolecular forces; increases with atomic mass or molar mass.

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